Chemistry

Math In Chemistry
Lesson 2
Lesson Plan
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Math In Chemistry – How to Measure

Objectives: Define SI, metric prefixes, estimate measurements, scientific notation

- Math – the language of Science
- How would you measure without a ruler?
- 1795 the French adopted the Metric System which has become the Susteme Internaionale d'Unites – SI
- Standards are kept at the International Bureau of Weights and Measures in Sevres, France and in the National Institute of Science and Technology (NIST) in Gaithersburg Maryland
- Base Quantities are length, time, mass
    - Length – meter
        - 1/10,000,000 distance from the north pole to the equator
        - Distance between two lines on a platinum-iridium bar
        - Distance light travels in a vacuum in 1/299,792,458 s
    - Time – second
        - 1/86,400 of a mean solar day
        - frequency of cesium-133 atom
    - Mass – kilogram
        - Platinum-iridium alloy cylinder
- Derived units are combinations of the base units.
    - m/sec for speed, $kg \cdot m^2/s^2$
- Which is more accurate – metric or English system?
    - Accuracy is the same, metric is based on 10 making it easier to use
- SI Prefixes

Prefix	Symbol	Multiplier	Scientific Notation	Example
femto	f	1/1,000,000,000,000,000	10^{-15}	femtosecond (fs)
pico	p	1/1,000,000,000,000	10^{-12}	picometer (pm)
nano	n	1/1,000,000,000	10^{-9}	nanometer (nm)
micro	µ	1/1,000,000	10^{-6}	microfarad (µF)
milli	m	1/1,000	10^{-3}	millimeter (mm)
centi	c	1/100	10^{-2}	centimeter (cm)
deci	D	1/10	10^{-1}	Deciliter (dL)
kilo	K	1000	10^{3}	Kilometer (km)
mega	M	1,000,000	10^{6}	Megabyte (Mb)
giga	G	1,000,000,000	10^{9}	Gigawatt (Gw)
tera	T	1,000,000,000,000	10^{12}	Terabyte (Tb)

-Examples

- 10 mm = 1cm

- 10 cm = 1 dm
- 10 dm = 1 m
- 1000 m = 1 km

- Scientific Notation
  - Convenient way to express very large or very small numbers
  - M x 10^n where $1 \leqslant M < 10$ and is multiplied by a whole number power of 10
  - Moving the decimal left, add to n. 1000. = 1000 x $10^0 \rightarrow$ 100 x $10^1 \rightarrow$ 10 x $10^2 \rightarrow$ 1 x 10^3
  - Moving the decimal right, subtract from n. .0001 $\rightarrow$ .0001 x $10^0 \rightarrow$ .001 x $10^{-1} \rightarrow$ .01 x $10^{-2} \rightarrow$ .1 x $10^{-3} \rightarrow$ 1 x 10^{-4}
  - Avg. distance from the sun to Mars is 227,800,000,000 m $\rightarrow$ 2.278 x 10^{11} m
  - The mass of an electron is 0.000,000,000,000,000,000,000,000,000,000,911 kg $\rightarrow$ 9.11 x 10^{-31} kg
  - Some calculators show scientific notation as ME^n , students should always write answer in full scientific notation (M x 10^n)

- Addition and Subtraction Using Scientific Notation
  - If numbers have the same exponent *n*, then add or subtract the values of M leaving *n* the same.
  - 4 x 10^8 m + 3 x 10^8 m = 7 x 10^8 m
  - 4.1 x 10^{-6} kg – 3.0 x 10^{-6} kg = 1.1 x 10^{-6} kg
  - 4.01 x 10^6 m +1.89 x 10^2 m = 4.01 x 10^6 m - .000189 x 10^6 m = 4.01 x 10^6 m

- Multiplication and Division Using Scientific Notation
  - Multiply the value and units of M, add exponents *n*.
  - Divide the values and units of M, subtract the exponent *n* of the divisor from the exponent *n* of the dividend
  - (4 x10^3 kg) (5 x 10^{11} m)
    4 x 5 = 20, kg x m = kg•m, 3 + 11 = 14
    20 x 10^{14} kg•m = 2 x 10^{15} kg•m
  - (8 x 10^6 m^3)/(2 x 10^{-3} m^2)
    8/2 = 4, 6-(-3) = 9, 3-2 = 1
    4 x 10^9 m

- Converting Units – Factor-Label Method
  - Convert 465g $\rightarrow$ kg
  - Setup a conversion factor of 1. Knowing that 1 kg = 1000g then we can construct 1 = 1 kg/1000 g or 1 = 1000g/1 kg
  - Multiplying or dividing by 1 does not change a value

- Set up the conversion such that units cancel out:
- $465g = 465g\left(\frac{1kg}{1000g}\right) = \frac{465g \times 1kg}{1000g} = \frac{465kg}{1000} = 0.465kg$
- If units do not work out, check your conversion factor

- Example Problems
  - 1.1 cm to meters
  - $1.1cm = 1.1cm\left(\frac{1m}{100cm}\right) = \frac{1.1cm \times 1m}{100cm} = \frac{1.1m}{100} = \frac{1.1m}{10^2} = 1.1x10^{-2}m$
  - or $1.1cm = 1.1cm\left(\frac{1m}{10^2 cm}\right) = \frac{1.1cm \times 1m}{10^2 cm} = \frac{1.1m}{10^2} = 1.1x10^{-2}m$
  - or $1.1cm = 1.1cm\left(\frac{10^2 m}{1cm}\right) = \frac{1.1cm \times 10^2 m}{1cm} = \frac{1.1x10^2 m}{1} = 1.1x10^{-2}m$

  - 76.2 pm to mm
  - $76.2pm = 76.2pm\left(\frac{1m}{10^{12} pm}\right)\left(\frac{10^3 mm}{1m}\right) = \frac{76.2pm \times 1m \times 10^3 mm}{10^{12} pm \times 1m} =$
    $\frac{76.2 \times 10^3 mm}{10^{12}} = 76.2 \times 10^{-9} mm = 7.62x10^{-8} mm$

  - $76.2pm = 76.2pm\left(\frac{1mm}{10^9 pm}\right) = \frac{76.2 \times 1mm}{10^9} = \frac{76.2mm}{10^9} =$
    $76.2x10^{-9} mm = 7.62x10^{-8} mm$

- Precision and Accuracy
  - Precision – the exactness of measurement (tolerance)
    - How close do all the measurements to each other
  - Accuracy – how well the results agree with a standard value
    - Instrument must be calibrated to known standard
- Qualitative Measurement vs. Quantitative Measurement
  - Qualitative – descriptive, non-numerical
    - You feel hot – you might have a fever
    - The ice cream is very cold
  - Quantitative – numerical, usually with units
    - The thermometer indicates that you temperature is 39.2°C – you have a fever
    - Ice cream at -21°C is hard and cold
- Error
  - Error = accepted value – experimental value
    - Can be positive or negative

- Percent Error
    - Percent error =(|error|/accepted value) x 100%
- Significant Digits – non-zero numbers
    - Draw diagram of mm ruler, 8.64 bar and a cm ruler
    - Note that the bar is 8.6 mm + a little by the mm ruler
    - The little is estimated to be .4 mm
    - So the length of the bar is 8.64 mm by the mm ruler – 8.6 can be seen, the .04 is an estimate
    - Note that the bar is 8 + a little by the cm ruler
    - The little is estimated to be .6 cm
    - So the length of the bar is 8.6 cm by the cm ruler – 8. can be seen, the .6 is an estimate
    - 8.64 is 3 significant figures, 8.6 is 2 significant figures
    - Redraw to have the bar 8.60 in length – 8.60 is 3 significant figures
- Significant Digits – zero's
    - Not all 0's are significant
    - Place holders are not significant
        - 0.0086 = 2 significant figures
        - 0.00860 = 3 significant figures
        - 186,000 = 3, maybe six – unknown since the decimal is not shown
        - 186 km is 3sf, 186.000 is 6sf
        - 1.86×10^5 has 3 sf
- Rules for Significant Digits
    - Non-zero digits are always significant
    - All final 0's after the decimal point are significant
        - These are 0's after the final non-zero digit
    - Zero between two other significant digits are always significant
    - Zeros used solely as place holders are not significant
- The following examples have 3 significant digits
    - 245 m   18.0 g 308 km   0.00623 g

- Arithmetic with Significant Digits
    - An answer can never be more precise that the least precise number
- Addition & Subtraction
    - Add 24.615 + 3.21 + 6.964 = 34.789
        - Since 3.21 has the least number of digits to the right of the decimal, the correct answer is 34.79
        - Same principle for subtraction
- Multiplication & Division
    - 3.22 cm x 2.1 cm = 6.762 cm^2 ->6.8 cm^2
    - Answer is rounded off to have the same number of significant digits as the factor with the least number of significant digits
    - 36.5 m / 3.414s = 10.691 m/s -> 10.7 m/s